

Total No. of printed pages = 5

Sc-103/Ch-I/1st Sem/2016/N

CHEMISTRY – I

Full Marks – 70

Pass Marks – 21

Time – Three hours

The figures in the margin indicate full marks for the questions.

Answer question No.1 and any six from the rest.

1. Fill in the blanks : $1 \times 10 = 10$

(i) E.C.E of Ag is _____ .

(ii) Conjugate acid of SO_4^{2-} is _____ .

(iii) Isotones are atoms of different atomic number but same number of _____ .

(iv) The oxidation number of Fe in Fe_3O_4 is _____ .

(v) The energy of first Bohr's orbit is _____ erg / atom.

[Turn over

(vi) 0.5 mole of oxygen occupies _____ ml volume at NTP.

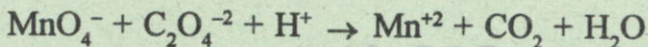
(vii) The vapour density of SO_2 gas is _____.

(viii) Number of protons present in O^{2-} is _____.

(ix) _____ is an example of acidic oxide.

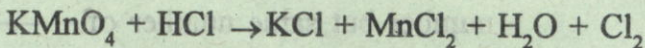
(x) The equivalent weight of CaCO_3 is _____.

2. (a) Balance the following equation by ion-exchange method : 3



Or

Balance the following equation by partial method :



(b) What do you mean by Lewis acid ? Give example. 2

(c) State Le Chatellier's principle and describe the effect of pressure, temperature and concentration in the manufacture of H_2SO_4 acid by contact process. 2+3=5

3. (a) State and explain Faraday's first law of electrolysis. 1+2=3
- (b) What do you mean by basicity of an acid and acidity of a base ? 2
- (c) What do you mean by decinormal solution ?
What volume of 0.5(N) NaOH is required to neutralize 50 ml of 1.5(N) HCl solution ?
1+3=4
- (d) What is indicator ? 1
4. (a) What are postulates of Bohr's atomic model ? 3
- (b) Prove that $M = 2D$ 3
- (c) Prove that $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$ 3
- (d) What is Pauli's exclusion principle ? 1
5. (a) What is modern periodic law ? 1
- (b) Derive De-Broglie's equation. 3
- (c) Write down the characteristics of transition elements. 3

(d) Calculate the pH of 0.0001(M) NaOH solution. 3

Or

Calculate the E.C.E of Zn (Atomic wt of Zn = 65).

6. (a) Calculate the number of moles and molecules present in 100 ml of CO_2 gas at NTP. 4

(b) Write down three industrial applications of catalyst. 3

(c) Write down the estimation of hardness of water by EDTA method. 3

7. (a) What do you mean by hydrolysis ? Why an aqueous solution of Na_2CO_3 is alkaline ?

2+2=4

(b) Write down the electron dot structure of N_2 molecule. 2

(c) How does ionisation energy vary in a period and in a group ? 2

(d) Identify the hardness causing salts among the following : 2

NaHCO_3 , $\text{Ca}(\text{HCO}_3)_2$, $\text{Al}_2(\text{SO}_4)_3$, MgSO_4 ,
 NH_4Cl , Na_2CO_3 .

8. Write short notes on any *four* : $2\frac{1}{2} \times 4 = 10$

(i) Common ion effect

(ii) Quantum numbers

(iii) Heissenburg's uncertainty principle

(iv) Solubility product

(v) Electrovalency.